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20 4 2-) Example complex N O [Cr(NH 2CH 2CH 2NH 2) 3] 3+
[Cr(C 2O 4) 3] 3- Multidentate ligands – ligands which form more than two co-ordinate bonds to a metal ion Ligand EDTA⁴⁻ forms 6 bonds porphyrin forms 4 bonds How it bonds Example
[Cu(EDTA)]²⁻ haemoglobin CH₂ CH₂ H₂N: :NH₂-O: C:O-O O NH₂
Cr NH₂ H₂N NH₂ H₂NH₂ 3+ O Cr NH₂ O O O O C ...

Chemsheets A2 1027 (Transition Metals)

Decomposition chemical reaction is the reaction where only one compound decomposes and results in two or more than two products. Pb(NO₃)₂ → PbO + NO₂ + O₂. In this equation, lead nitrate is being decomposed, which breaks down to form nitrogen dioxide, oxygen, and lead oxide. This is an example of a decomposition reaction.

49 Balancing Chemical Equations Worksheets [with Answers]

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Example 3 50.0 cm³ of 0.500 mol dm⁻³ copper sulphate solution was put in a calorimeter and 2.0 g of zinc powder added. The temperature of the solution rose by 24.0 C.

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Full worked solutions in: Chemsheets AS029 (Calorimetry) TASK 2 – Calorimetry calculations 1 -3800 kJ mol⁻¹ 2 -51.8 kJ mol⁻¹ 3 -2500 kJ mol⁻¹ 4 +20.9 kJ mol⁻¹ 5 -4203 kJ mol⁻¹ 6 -3920 kJ mol⁻¹ 7 -54.3 kJ mol⁻¹ 8 -1145 kJ mol⁻¹ 9 -1223 kJ mol⁻¹ 10 -154 kJ mol⁻¹ 11 exo, exo, endo 12 -503 kJ mol⁻¹

Chemsheets AS 029 (Thermodynamics) ANS 2

n is called the order of reaction with respect to that reactant,

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and is usually 2, 1 or 0 Order Zero order First order Second order
n 0 1 2 The effect the reaction rate is proportional the reaction
rate is proportional to $[A]^0$ (= 1) and so the rate is not affected
by changes in $[A]$ to $[A]^2$ $[A] \times 2$ rate unchanged rate $\times 2^2$
(i.e. $\times 4$)

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Page 7 Example 2 - butane 8 7 CH 6 5 4 3 2 1 0

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2,2-dimethylamine 2-methylbut-1-ene 3-methylbutan-1-ol
4-hydroxybutanone 3-methylbutanoic acid butyl methanoate
ethanenitrile propanal Decide which spectrum belongs to which
compound and draw the molecule next to the spectrum.

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Amount Of Substance Chemsheets

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SECTION 1 1H NMR - Weebly

Example 1 Imagine a reaction where R reacts with S to make
some products: $R + S \rightarrow$ products If we compare experiments 1
and 2: Initial $[R]$ Experiment mol drn-3 1.0 2.0 Init-ial $[S]$ mo drn
9.0 1.0 '2.0 The

Chemsheets Kinetics 1 Answers

the value of $-114.2 \text{ kJ mol}^{-1}$ in the first equation means that
 114.2 kJ of heat energy is released when 1 mole of H_2SO_4
reacts with 2 moles of NaOH . the value of $-57.1 \text{ kJ mol}^{-1}$ in the

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second equation means that 57.1 kJ of heat energy is released when $\frac{1}{2}$ mole of H_2SO_4 reacts with 1 mole of NaOH.

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